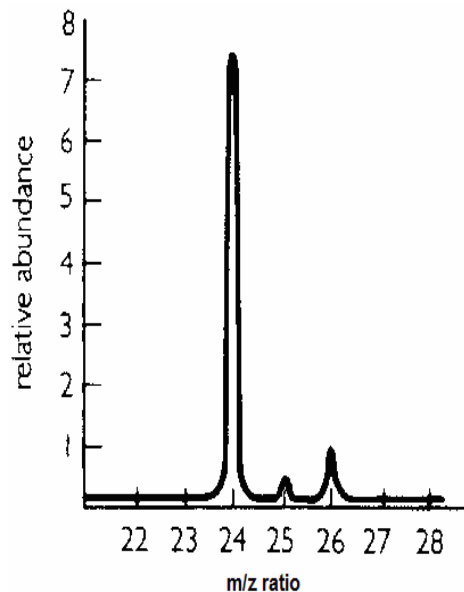


## Average Atomic Masses Homework

- 1) The output from a mass spectrometer is shown below. From the information provided, find the average atomic mass of the element being studied.



- 2) The three naturally-occurring isotopes of uranium are shown in the table below. Using this information, find the average atomic mass of uranium.

Isotope	Isotopic Mass (amu)	Abundance (%)
$^{234}\text{U}$	234.041	$5.5550 \times 10^{-3}$
$^{235}\text{U}$	235.044	0.7201
$^{238}\text{U}$	238.051	99.275

- 3) Explain how a mass spectrometer separates the isotopes of an element.