

Types of Intermolecular Forces

What is the strongest intermolecular force present for each of the following compounds?

- 1) water _____
- 2) carbon tetrachloride _____
- 3) ammonia _____
- 4) carbon dioxide _____
- 5) phosphorus trichloride _____
- 6) nitrogen _____
- 7) ethane (C₂H₆) _____
- 8) acetone (CH₂O) _____
- 9) methanol (CH₃OH) _____
- 10) borane (BH₃) _____

Types of Intermolecular Forces - Solutions

What is the strongest intermolecular force present for each of the following compounds?

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|--|---------------------------------|
| 1) water | hydrogen bonding |
| 2) carbon tetrachloride | London dispersion forces |
| 3) ammonia | hydrogen bonding |
| 4) carbon dioxide | London dispersion forces |
| 5) phosphorus trichloride | dipole-dipole forces |
| 6) nitrogen | London dispersion forces |
| 7) ethane (C ₂ H ₆) | London dispersion forces |
| 8) acetone (CH ₂ O) | dipole-dipole forces |
| 9) methanol (CH ₃ OH) | hydrogen bonding |
| 10) borane (BH ₃) | dipole-dipole forces |